5. Now pour back the first standard solution into its test tube (keep all your solutions! You may need to check your readings), then drain the cuvette upside-down onto a piece of tissue until excess liquid has drained away.
6. Partly fill the same cuvette with the next standard solution, and measure the absorbance of that solution.
7. Repeat this procedure: adding 1 mL, measure, pour back, drain with all the other standard solutions. Checking at intervals that the distilled water cuvette reads zero absorbance.
8. Take the absorbance reading for each standard in turn, and record all your readings to 3 decimal places in the table below copied into your Workbook.

Standard	1	2	3	4	5	6	7
Concentration (M)	0.005	0.01	0.02	0.04	0.06	0.08	0.10
Absorbance at 800 nm	0.038	0.082	0.241	0.433	0.728	0.965	1.205

Your values should rise in a regular ascending series. If they do not, recheck each absorbance value from the samples you have poured back into the test tubes.

9. Finally measure the absorbance of the five control and five experimental samples (do not dilute) provided (they are all different) and record the reading in the table below, to determine the amount of cupric sulphate in each ' test sample' .

Control	Unaffected farm Sample 1	Unaffected farm Sample 2	Unaffected farm Sample 3	Unaffected farm Sample 4	Unaffected farm Sample 5
Absorbance at 800 nm	0.365	0.305	0.187	0.307	0.247
Experimental	Affected farm Sample 1	Affected farm Sample 2	Affected farm Sample 3	Affected farm Sample 4	Affected farm Sample 5
Absorbance at 800 nm	0.424	0.368	0.604	0.492	0.499