

No pre-lab next
WEEK

Cool off for 5 minutes
remeasure it: 25.52g
25.512
- 25.36

25.55

Finish
25.52g

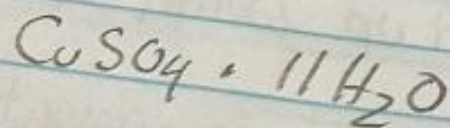
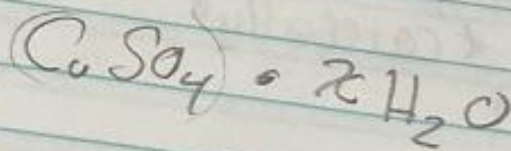
$$\frac{.16g \text{ CuSO}_4}{159.609 \frac{\text{mol}}{\text{g}}} =$$

$$\frac{159.609 \text{ g/mol CuSO}_4}{100 \times 10^{-3}} =$$

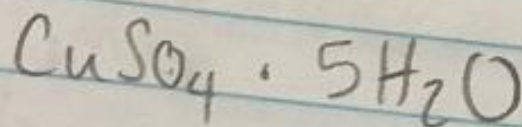
$$\frac{25.36}{25.55} = 0.99$$

$$119 \text{ g H}_2\text{O}$$

$$\frac{18.01528 \text{ g H}_2\text{O}}{1.05 \times 10^{-2}} =$$



should have been



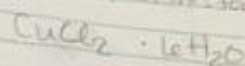
$$\frac{1.05 \times 10^{-2} \text{ moles H}_2\text{O}}{1.00 \times 10^{-3} \text{ moles CuSO}_4} = 10.5$$

Copper Sulfate X hydrate 100 Cu 17mg Acid Zn 100 Base

Gravimetric analysis method uses weight to quantitatively determine the composition of a solid analyte.
anhydrous "without water" weights initial vs after determining water content
Hydrates are compounds that are salt with water molecules
chemically bonded to them

solving for x: $CuCl_2 \cdot xH_2O$

measured in water form



ring stand + ring clamp
Clay Triangle - rest on the ring
Crucible rest on the clay triangle
↳ top slightly on top of crucible (2 blue caps)
heat crucible but before weigh the crucible + top before hand mass
handling use the tongs * carefully *
Crucible goes on Hotpad

Heat it up weigh the heat it up weigh if its dripping you
Still have moisture.

crucible w/ top: 25.36g

equipment

weight boat
Bunsen burner

Burner striker

ring stand w/ clamp

clay triangle

tongs

crucible w/ top

unknown hydrate compounds
hot pad.

balance / tare weight boat put

place crucible on clay triangle w/ tongs

bunsen burner on 5 minutes

took off to add hydrate.

place back on clay triangle heat up for additional 5 minutes

Took off using the tongs + place on hot pad additional 5 minutes

remeasured on digital balance: 25.55g 1st

reheat additional 5 minutes

remeasured on digital balance: 25.52g 2nd

reheat additional 5 minutes

color when heated
Blue changed to white